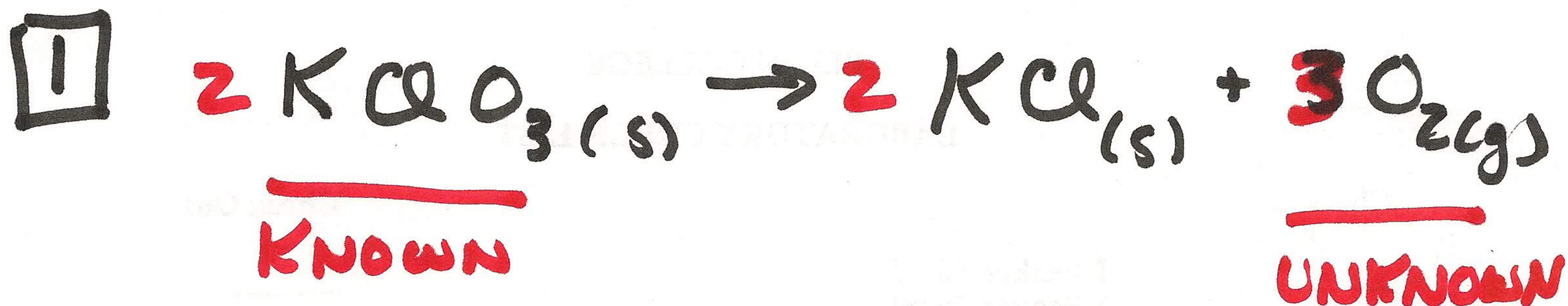


SOLUTIONS 4

4.1



$$1 \times 39.1 = 39.1$$

$$1 \times 35.5 = 35.5$$

$$3 \times 16 = 48$$

$$\underline{122.6 \text{ g/mole}}$$

$$32 \text{ g/mole}$$

KNOWN:

$$6.09 \text{ g KClO}_3 \left[\frac{1 \text{ mole KClO}_3}{122.6 \text{ g KClO}_3} \right]$$

$$= 0.0497 \text{ moles of KClO}_3$$

MULTIPLY BY STOICHIOMETRIC RATIO:

$$0.0497 \text{ moles KClO}_3 \left[\frac{3 \text{ moles O}_2}{2 \text{ moles KClO}_3} \right]$$

$$= \underline{0.0745 \text{ moles O}_2}$$

