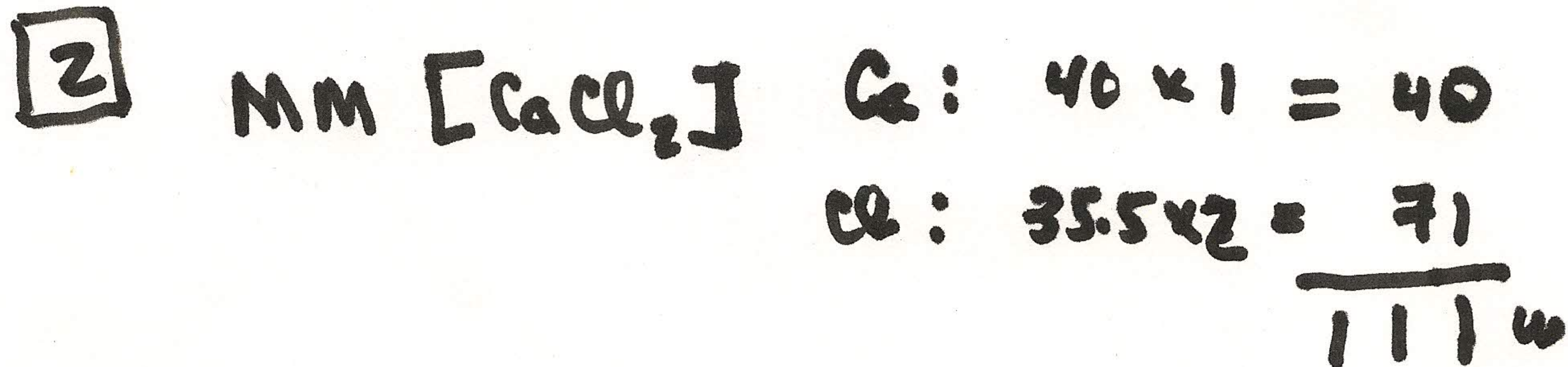


CHAPTER TWO - SOLUTION

2.1

1 ELECTRONS



% COMPOSITION

$$\frac{71 \text{ u}}{111 \text{ u}} = \frac{\text{MASS ELEMENT}}{\text{MOLAR MASS}} \times 100 = 64\%$$

MASS OF CHLORINE

$$59.7 \text{ g } [CaCl_2] \times 0.64 = 43.2 \text{ g OF CHLORINE}$$

